Chemistry Acid Base Equilibrium Review Crossword

seven/pH of a neutral substance

water/example of a neutral substance

less/acids have a pH \_\_\_\_ than 7

more/bases have a pH \_\_\_\_ than 7

bitter/how bases taste

sour/how acids taste

base/slippery

base/cleaners

acid/food

blue/bases turn litmus paper \_\_\_\_\_

red/acids turn litmus paper \_\_\_\_\_

hydroxide/bases has more \_\_\_\_\_\_\_\_\_ ions

hydrogen/acids have more \_\_\_\_\_\_\_\_\_ ions

acid/acid or base: HNO3

base/acid or base: NaOH

electrolytes/acids and bases both conduct electricity in solution, so we call them \_\_\_\_\_\_\_\_\_\_\_\_\_

twelve/which is a stronger base, pH of 11 or 12?

salt/acid base neutralization always produces water and a \_\_\_\_\_\_\_

doublereplacement/acid base neutralization is always this type of reaction

acid/what would you add to a base to neutralize it?

acid/reacts with metals

base/reacts with fats and oils

saturated/solution with the maximum amount of solute dissolved

unsaturated/solution with less than the maximum amount of solute dissolved

singlereplacement/type of reaction: A + BC -> B + AC

synthesis/type of reaction: A + B -> AB

reactants/things you start with in a chemical reaction

products/things you end with in a chemical reaction

products/this is favored (a higher concentration exists at equilibrium) if Kw is above 1.

Reactants/this is favored (a higher concentration exists at equilibrium) if Kw is below 1.

Left/In the reaction H2 + N2 -> NH3, removing H2 would shift equilibrium to the \_\_\_\_\_\_\_

right/In the reaction H2 + N2 -> NH3, adding N2 would shift equilibrium to the \_\_\_\_\_\_\_

Left/In the reaction H2 + N2 -> NH3 + heat, adding heat energy would shift equilibrium to the \_\_\_\_\_\_\_

Base/after proton transfer occurs, the acid forms the conjugate \_\_\_\_\_

Gas/when determining equilibrium based on pressure, you must county up to moles of \_\_\_\_ on each side of the equation

Less/When you raise the pressure, equilibrium shift towards the side with \_\_\_\_\_\_ gas

Acceptor/Bronsted-Lowery defined a base as a proton \_\_\_\_\_\_\_\_\_\_\_

Donor/Bronsted-Lowery defined an acid as a proton \_\_\_\_\_\_\_\_\_\_\_