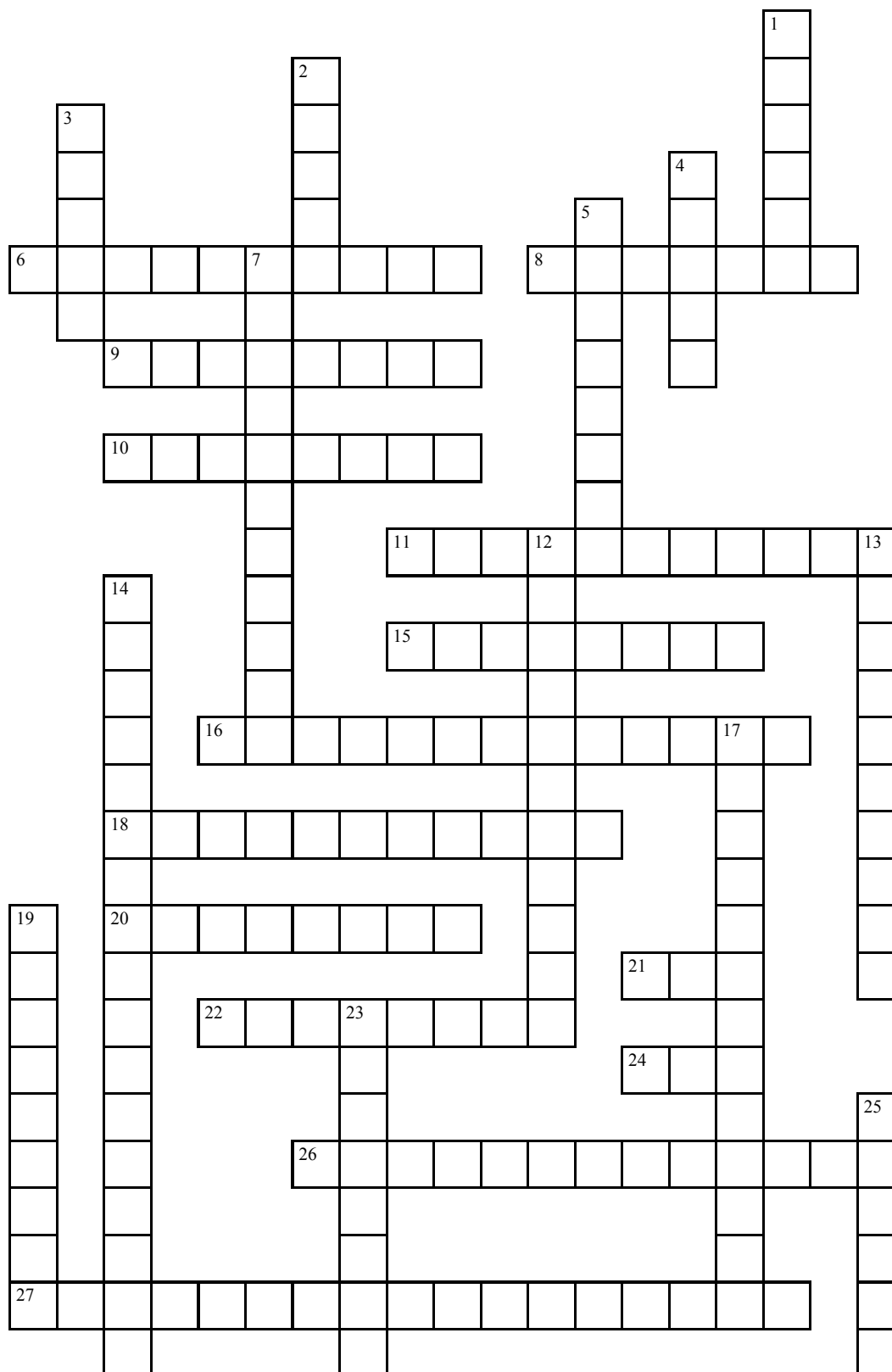# Reactions Review: 11/6

ACROSS

6 Tell the Reaction Type: C2H4 + O2 -> CO2 + H2O

DOWN

1 When you heat objects, the atoms move faster causing the object to _____

8 Reactions rates are based on the idea that for chemicals to react with each other the atoms must _____

9 A burning splint would reignite in the presence of this gas

10 Lowers activation energy to increase rate of a reaction

11 You could speed up a reaction by raising the _____

15 An exothermic reaction occurs when energy is _____

16 Lime water turns milky white in the presence of this gas

18 Beaker gets cold during reaction

20 Another word for heat energy

21 Where are the most reactive metals on the activity series?

22 In an exothermic reaction, the delta-H (change in enthalpy is _____

24 Indicator of chemical change when you see bubbles: production of a ___

26 5 mentos react faster than 2. This is because we raised the _____

27 Tell the Reaction Type: K + NaCl -> Na + KCl

2 In an exothermic reaction, heat is listed on which side of the yield sign?

3 Indicator of chemical change evident during rusting: a change in _____

4 Indicator of a chemical change when you see a precipitate

5 In an endothermic reaction, the delta-H (change in enthalpy) is _____

7 You could increase the rate of a reaction by breaking up the chemicals into smaller pieces (increase _____ _____)

12 Term for a solid that forms out of aqueous solution, making the solution look cloudy

13 Beaker gets hot during reaction

14 Tell the Reaction Type: NaOH + HBr -> H2O + NaBr

17 Tell the Reaction Type: H2O2 -> H2 + O2

19 Tell the Reaction Type: H2SO3 + O2 -> H2SO4

23 An endothermic reaction occurs when energy is _____

25 Indicator of a chemical change: when _____ is gained or released