Atomic structure review

lead/Pb

gold/Au

negative/charge of an electron

tin/Sn

eighteen/numbers of electrons that fit in the 3rd principle energy level

Iron/Fe

neutral/overall charge of an atom

isotopes/atoms of the same element with different numbers of neutrons

homogeneous/solutions are an example of this type of mixture

protonsneutrons/two particles you add together to get an atoms mass

molecule/smallest particle of a compound you can have

potassium/K

oneamu/mass of a proton

electron/subatomic particle with almost no mass

protons/an atoms atomic number is equal to its \_\_\_\_\_\_\_\_

sodium/Na

bohr/developed the planetary model of the atom

amu/One Twelfth the mass of a Carbon-12 atom is the definition of an \_\_\_\_\_\_

atom/smallest particle of an element you can have

chemical/change in matter where the composition of the matter changes (ex:rusting)

physical/change in matter where the composition of the matter does not changes (Ex: dissolving in water or changing state)

sublimation/solid turns straight to gas

condensation/gas turns to liquid

valence/term for electrons in the outermost shell

group/column in the periodic table

period/row in the periodic table

halogens/most reactive nonmetals

alkalimetals/most reactive metals

noblegases/least reactive elements: they are the only ones found uncombined in nature

metalloids/elements that touch the staircase

transitionmetals/another name for the “B” Groups

liquid/state of matter with a definite volume but indefinite shape

gas/the only state of matter than can be compressed (indefinite volume)

electron/type of microscope that can be used to see atoms

eight/atomic number of oxygen

ten/number of neutrons in oxygen-18

volume/term for the space an object takes up

silicon/element in Group 4A, Period 3